



Canadian Advances in Thermochemical H₂ Production in the Context of Conventional Electrolysis

**Alistair I. Miller
Romney B. Duffey
Sam Suppiah**

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Canada 



AECL
Atomic Energy
of Canada Limited

EACL
Énergie atomique
du Canada limitée



Outline

The market for hydrogen

SMR vs LTE

The GIF context

AECL work on sulphur thermochemical cycles

Collaboration with USDOE on copper chloride cycles



Where will the demand be?

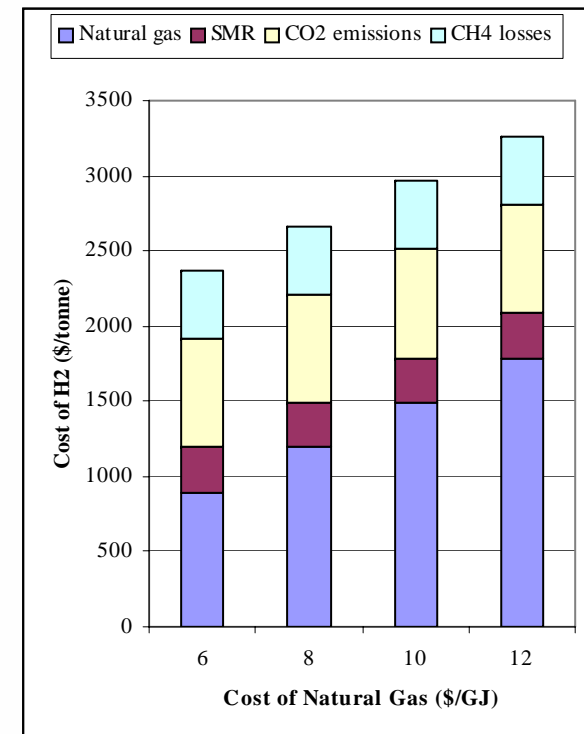
- Fuel for road vehicles?
 - Later perhaps but uncertain
 - Depends on battery vs fuel cell development
- More likely for larger vehicles (trains, ships)
- Big market is for upgrading petroleum
 - Exists and is growing rapidly
 - Especially in the oil sands developments in northern Alberta
 - Needs 3 to 5 kg H₂/bbl
 - Expect over 2 million bbl/d by 2015
 - 1 GWe = 160 000 bbl/d





How will H₂ be made?

- Conventionally come from natural gas by SMR
 - Cost has risen fast
 - Realistic to base on oil:gas at 6:1
 - Add 70 \$/t CO₂
 - Add 3% leakage of CH₄ from well to end use
 - Supply of natural gas is uncertain
 - All Mackenzie pipeline output could go to oil sands upgrading
- Need a new way
 - High-temperature thermochemical?
 - High-temperature electrolysis?
 - Conventional low-temperature electrolysis?





LTE will be available much sooner

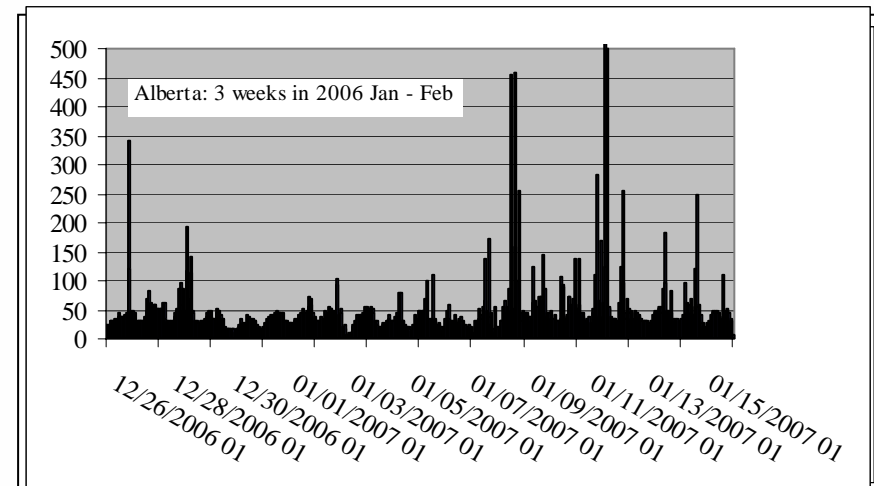
- Make it using Generation III+ reactors
 - Could be deployed by 2015
- Key is to produce H₂ with off-peak electricity



- Preferably with variable-current cells
- Needs large-scale storage
 - In salt caverns

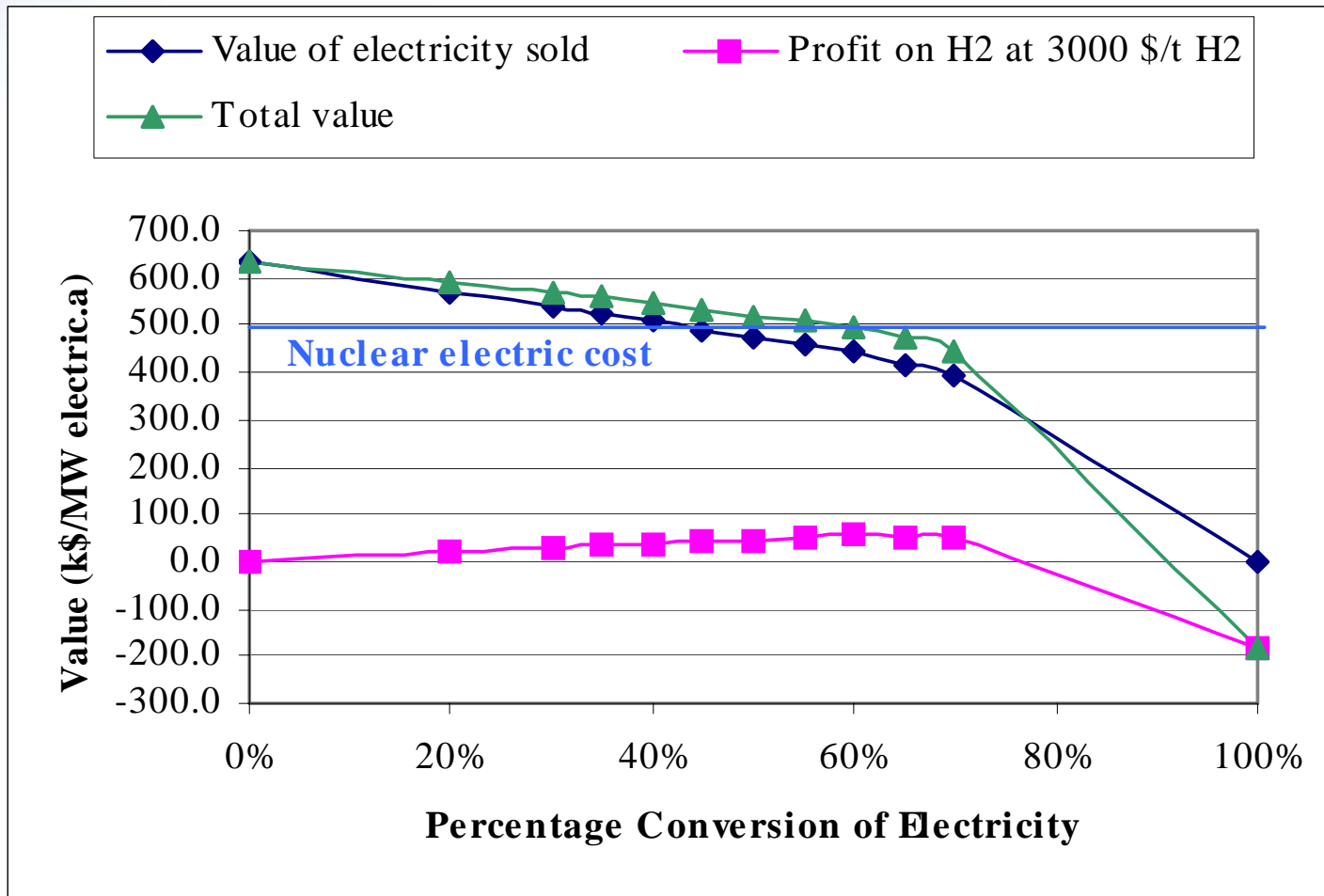
⇒ Alberta case

- 550 \$/kW cells
- 5000 \$/t storage
- Applying real-time Alberta power prices





Intermittent H₂ Production





And later?

Within the GenIV, Canada focuses on SCWR with crosslink to VHTR

<i>Acronym</i>	<i>Spectrum</i>	<i>Fuel cycle</i>
• SFR Sodium Cooled Fast Reactor	Fast	Closed
• LFR Lead Alloy Cooled Reactor	Fast	Closed
• GFR Gas Cooled Fast Reactor	Fast	Closed
• VHTR Very High Temperature Reactor	Thermal	Once-through
• SCWR Supercritical Water Cooled Reactor	Th. & F.	Once-t. & Closed
• MSR Molten Salt Reactor	Thermal	Closed

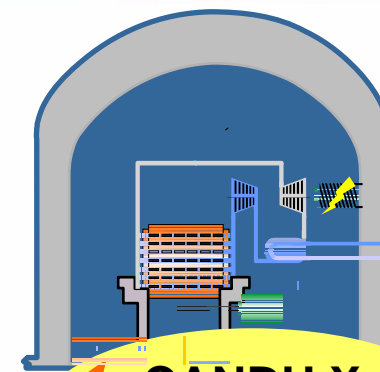




CANDU Evolution

**Advanced
CANDU Reactor**

<5 years



**CANDU X
(SCWR)**

20+ years

**Project
Experience**

**Operating
Feedback**

**Market
Pull**

**Technology
Push**

**Current
Generation CANDU**

Product Evolution

- Improved Economics
- Enhanced Safety
- Enhanced Operability



CANDU SCWR Concept

- Started in 1994 as Candu X Program
- Establish the design limits and ultimate potential
- Main CANDU features are retained.
 - Horizontal modular channels.
 - Heavy water moderator.
- Supercritical light water coolant (higher efficiency).
- Advanced fuel channel design (internal insulation without calandria tube).
- Options systematically studied
 - Mark 1: indirect cycle $T_{out} \sim 400^+ \text{ }^\circ\text{C}$ set by existing Zr
 - Mark 2: direct cycle $T_{out} \sim 600^+ \text{ }^\circ\text{C}$ set by existing turbine
 - Mark3: multiple cycle $T_{out} >850^+ \text{ }^\circ\text{C}$ set by known materials



Thermochemical Work in Canada
In collaboration within the GIF
and through I-NERI agreements
with the USDOE

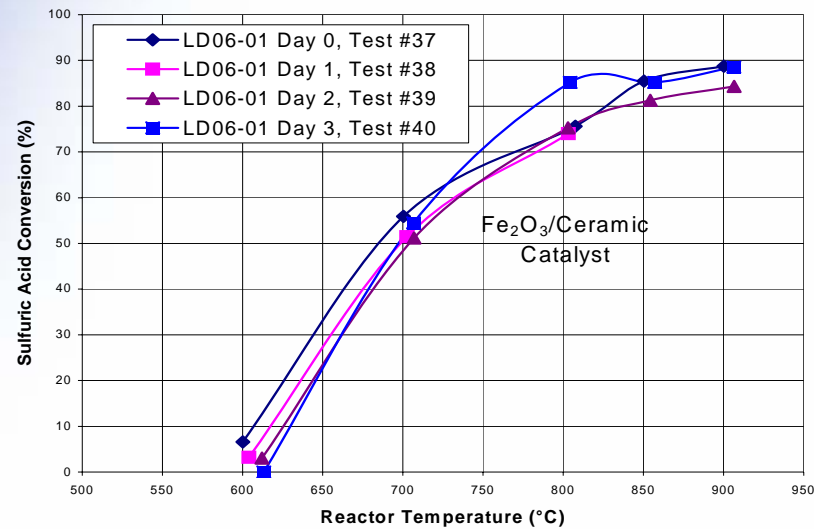


The H₂SO₄ Side of I/S and other S Cycles

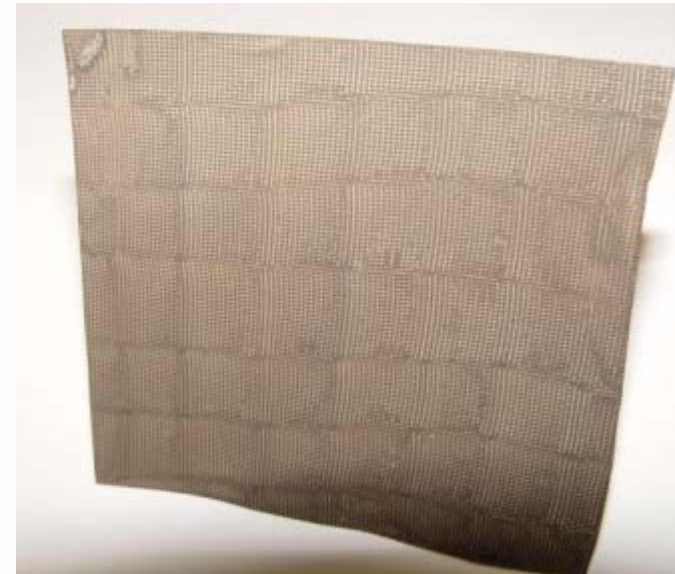
- $\text{H}_2\text{SO}_4 \rightarrow \text{SO}_3 + \text{H}_2\text{O}$
 - Majority of energy; lower temperature (< 500°C)
- $\text{SO}_3 \rightarrow \text{SO}_2 + \frac{1}{2} \text{O}_2$
 - Minority of energy; higher temperature (> 700°C)
 - Could avoid a high temperature reactor by providing direct electric heating of a substrate on which catalyst deposited
 - Work so far on selecting catalysts



Assessing catalysts for SO₂ decomp.

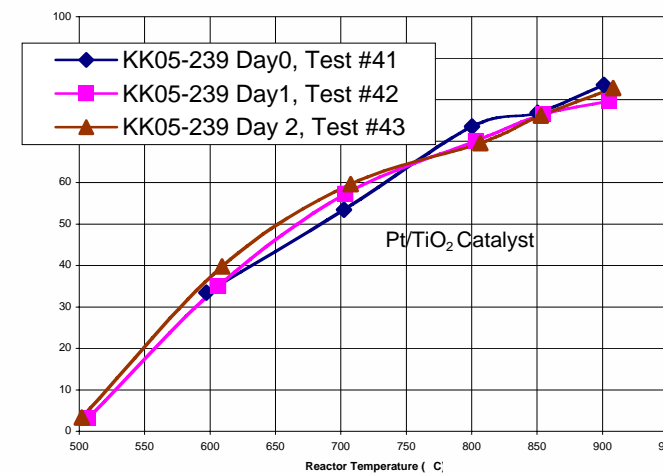


Fe catalyst



A metal (textured Inconel 800) sheet coated with catalyst for SO₃ decomposition

Pt catalyst





Copper chloride cycles

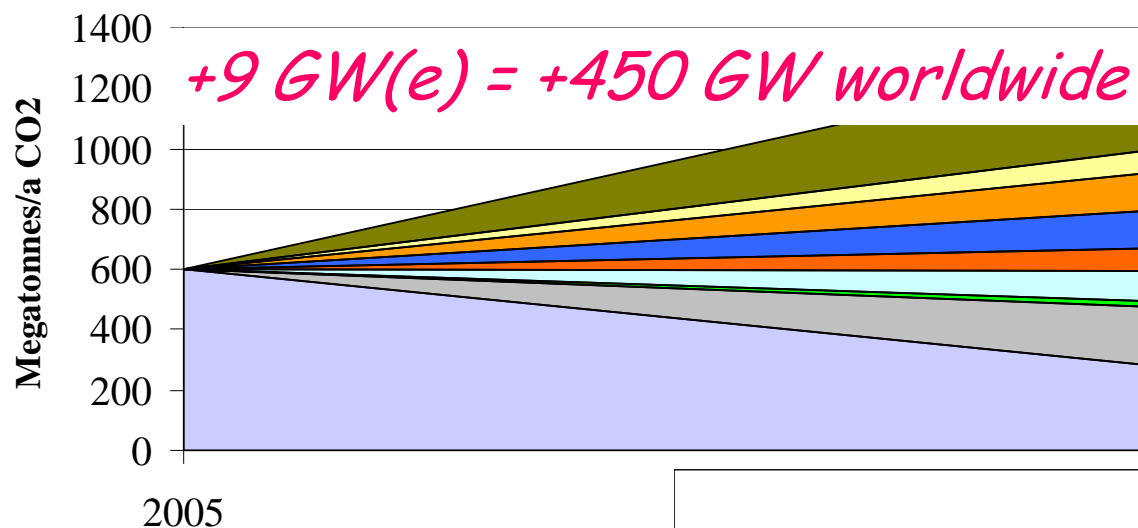
- Work led by USDOE at Argonne (Michelle Lewis)
- AECL is currently focused on the electrochemical step

#	Reaction Stoichiometry	Temperature (°C)
1	$2\text{Cu} + 2\text{HCl}(\text{g}) \rightarrow 2\text{CuCl}(\text{l}) + \text{H}_2(\text{g})$	425-450
2	$4\text{CuCl}(\text{s}) \rightarrow 2\text{CuCl}_2(\text{a}) + 2\text{Cu}$	<100
3	$2\text{CuCl}_2(\text{s}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{Cu}_2\text{OCl}_2(\text{s}) + 2\text{HCl}(\text{g})$	300-375
4	$\text{Cu}_2\text{OCl}_2(\text{s}) \rightarrow 2\text{CuCl}(\text{l}) + \frac{1}{2}\text{O}_2(\text{g})$	450-530

- Or a variant on reaction #2: $2\text{CuCl} + 2\text{HCl} \rightarrow 2\text{CuCl}_2 + \text{H}_2$
 - Avoids solid phase
 - Preliminary testing yields H_2 from both reactions at ~ 0.65 V

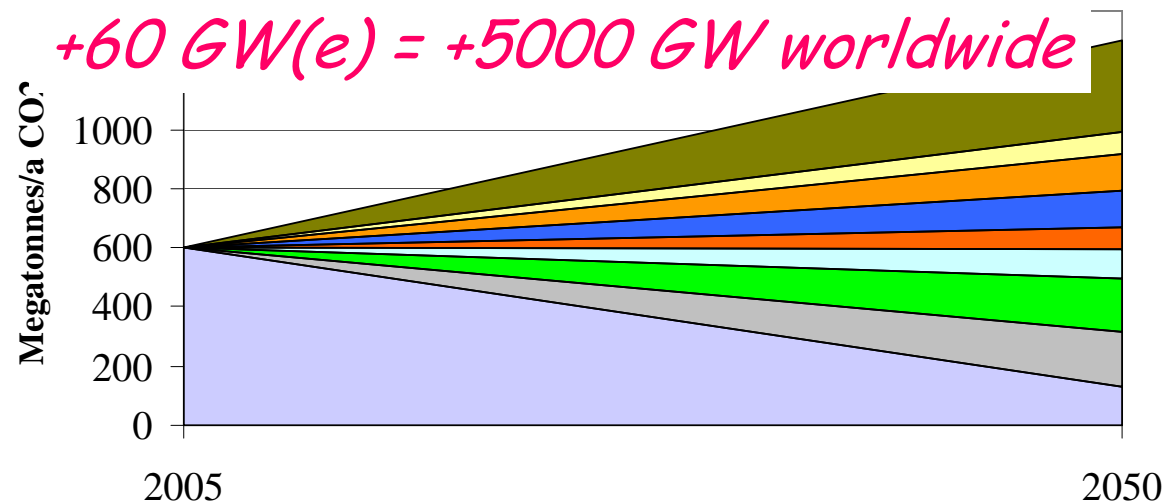


Placing Canada in Global Context



As in NRTEE:
⇐ 50% reduction

*NRTEE + large
nuclear deployment:
75% reduction ⇒*



Carbon emitting	Carbon capture and sequestration	Nuclear electricity
Biofuels and alternative fuels	Renewable electricity	Changes in energy intensity
Urban form	Co-generation	Energy eff & conservation

A large, stylized blue logo in the center of the page. It features a large letter 'A' with a smaller 'A' inside it, and a small figure of a person running to the right. To the right of the logo, the text 'AECL' is stacked above 'EACL' in a bold, serif font.

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